

LOWERING SYNTHESIS TEMPERATURE OF hBN BY IMPROVEMENT OF PRECURSOR

Erhan Budak^{1,2*}, Ramazan Lok², Ercan Yilmaz^{2,3}

¹Department of Chemistry, Faculty of Art and Science, Bolu Abant Izzet Baysal University, Bolu, Turkey

²Center for Nuclear Radiation Detector Research and Applications, Bolu Abant Izzet Baysal University, Bolu, Turkey

³Department of Physics, Faculty of Art and Science, Bolu Abant Izzet Baysal University, Bolu, Turkey

Abstract. In this study, hexagonal boron nitride (hBN) was synthesized with the modified O'Connor method in the presence of different additives. Structural properties of the synthesized materials were determined by X-Ray Diffraction, (XRD), Fourier Transform Infrared Spectroscopy, (FTIR), and Scanning Electron Microscopy (SEM). It was found that improving the precursor using different additives played a positive role by lowering the formation temperature. The average grain sizes (21–24 nm) and graphitization index (2.44–3.45) of hBN samples were calculated from the XRD pattern.

Keywords: Hexagonal boron nitride, Nanocrystalline materials, O'Connor method

1. INTRODUCTION

Boron nitride (BN) is an extraordinary structure in materials science. Due to special bonding behaviors of boron and nitrogen, BN can exist in many different structures. Graphite like hexagonal boron nitride (hBN), with a wide range of applications, such as lubricants, electrical insulators, refractories, catalyst supports, and laser devices [1] has been subjected to numerous studies in the last twenty years [2]-[4]. Although there are several different methods of hBN production, high temperature (≥ 1500 °C) solid-state synthesis reactions are most commonly used at industrial scale [5].

Due to simplicity and possibility to make changes in precursor preparation, the O'Connor method [6] was chosen to synthesize hBN in this study. In previous studies [2], [7], [8], it was claimed that the use of metal salts in the synthesis of hBN gave positive results: interlayer spacing was affected by an electronic interaction between the transition metal and hBN layers (d- π interaction) [9], and metal salts worked as a catalyst, increasing the surface area of boron oxide to complete reactions with ammonia [5].

In our previous studies [2], the addition of lithium oxide to the precursor yielded better result than other lithium salts added at 1100°C. In the current study, the effect of different metal oxides, namely MgO, CaO and BaO (Group II oxides), on the formation of hBN has been examined.

2. EXPERIMENTAL

BN samples were prepared using the O'Connor method [6]. The 2 g of B₂O₃, 4 g of CO(NH₂)₂ (urea) and 1.2 g of Group II oxides (MgO, CaO, and BaO) were ground in a mortar. Then the mixture was pre-heated

to 200°C for 2 h. After that, the obtained mixture (precursor) was pulverized in the mortar again. The precursor was heated to 1100°C at the heating rate of 5 °C/min in a tube furnace under NH₃ stream (flow rate 120 mL/min) for 2 h. Raw product was leached in a 10% HCl solution to remove all non-reacted products and dried in an oven at 100 °C.

The hBN stretching vibrations were determined using a Perkin Elmer Spectrum Two FTIR-ATR spectrophotometer. XRD patterns were obtained by a Rigaku Multiflex diffractometer using CuK α radiation, while micrographs were taken by the Jeol JSM 6390LV Microscope. For SEM images, the powder was placed on a tape and adhered onto a brass sample holder, followed by sputtering of gold.

3. RESULTS AND DISCUSSION

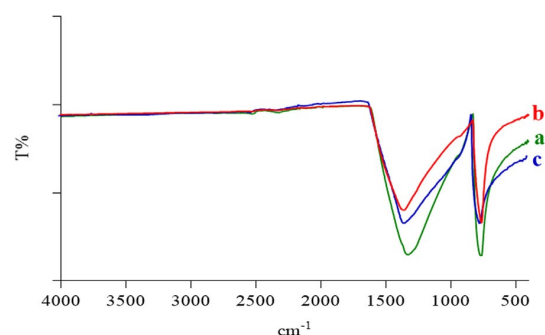


Figure 1. FTIR spectra of hBN samples at 1100 °C.
a) MgO/hBN, b) CaO/hBN, c) BaO/hBN.

Figure 1 shows two strong characteristic peaks, labeled as BN in-plane and out-of-plane vibrations, at ~ 1340 cm⁻¹ and ~ 761 cm⁻¹, respectively [7], [8], [10],

* erhan@ibu.edu.tr

[11]. There is no major difference of these spectra from those typical for industrially produced hBN, although the synthesis temperature is significantly lower; also, the absorption peaks are sharp.

Crystal structure of the hBN samples was examined by XRD and the results were matched with ICDD card No: 34-421. No different phases which could have originated from Group II oxides and boron oxide were found. In diffractograms in Figure 2, all peaks, namely (002), (100), (101), (102), (004), (103), (104), (110), (112), and (006), were labeled as hBN.

The separation of (100) and (101) peaks, accompanied by the sharpening of (002), the appearance of (004), and increasing intensity of (102) peaks, are the most significant proofs for the crystallinity of hBN [1], [12]. According to these concepts, we can state that the obtained hBN powders are well crystallized. In addition, the appearance of (104) (Figure 2c) shows that the degree of crystallization was very high.

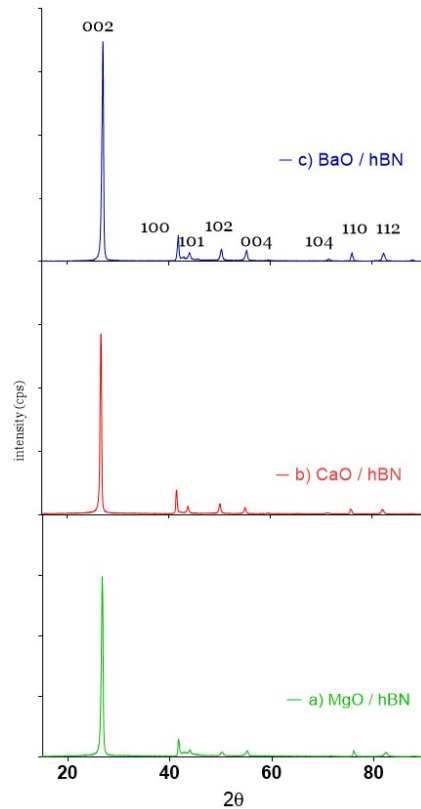


Figure 2. XRD pattern of hBN samples at 1100 °C. a) MgO/hBN, b) CaO/hBN, c) BaO/hBN.

The term “graphitization index” (G.I.) was used for estimating the degree of crystallization of the hBN and it was calculated as:

$$G. I. = \frac{\text{Area}[(100) + (101)]}{\text{Area}[(102)]} \quad (1)$$

G.I. value can be in the range from 1.6 to 50. A higher value of G.I. would mean less three-dimensional ordering (amorphous structure) in hBN and vice versa [12], [13]. In Table 1, the G.I. values of hBN are given; these low values indicate that the samples have high three-dimensional order in the crystal structure.

Table 1. Calculated lattice parameters, grain size, and graphitization index of hBN samples.

Metal Oxides	Lattice Parameter (Å)		Grain Size (nm)	Graphitization Index
	a _o	c _o		
BaO	a _o	2.506	24	2.44
	c _o	6.700		
	d	3.350		
CaO	a _o	2.496	22	2.74
	c _o	6.602		
	d	3.301		
MgO	a _o	2.502	21	3.45
	c _o	6.664		
	d	3.332		

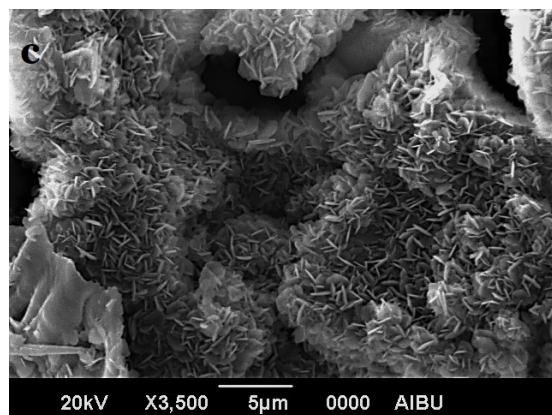
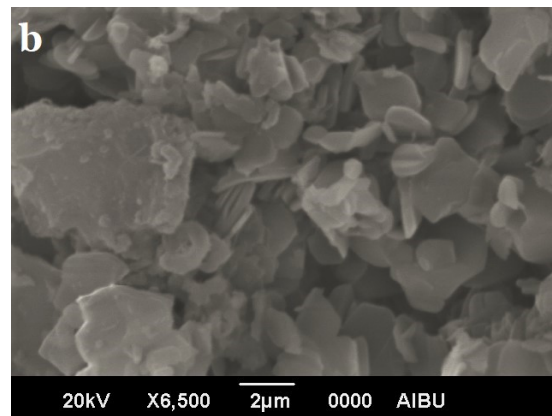
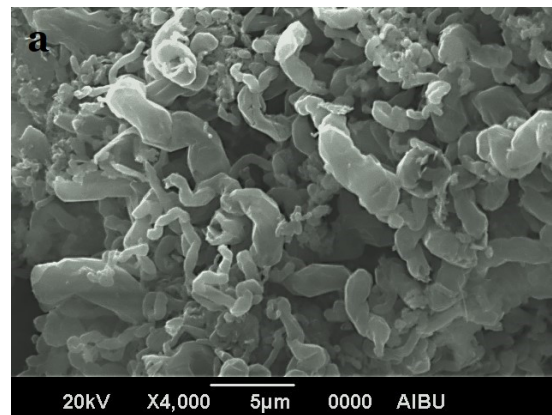


Figure 3. SEM images of hBN samples at 1100 °C. a) MgO/hBN, b) CaO/hBN, c) BaO/hBN.

The lattice constants, interlayer distance, and average grain size (Scherrer equations) were calculated for the samples after refinement and are also given in Table 1. The results are very close to the reported values for hBN; namely the ICDD card No: 34-421 has $a_0 = 2.504 \text{ \AA}$, $c_0 = 6.656 \text{ \AA}$, and $d = 3.328 \text{ \AA}$.

The use of both ammonia and metals lowered the formation temperature of hBN – metals act as catalysts [14]-[16] and ammonia acts as a nitrating agent [10], [17]. Furthermore, the usage of metal salts increases the boron oxide surface so that the full reaction of boron oxide and ammonia is possible [5]. According to these statements and results of the aforementioned FTIR and XRD analyses, Group II oxides play the role of a catalyst, lowering the formation temperature of hBN and increasing its crystallinity. The use of BaO provided best results among the three studied oxides; the reason may be its low melting point and greater atomic radius.

SEM images (Figure 3) indicate that the samples conglomerated and plate like crystals. The morphology difference of hBN possibly originates from the use of different metals, as in previously reported work [2], [18].

4. CONCLUSION

In summary, nano-sized hBN powders were synthesized at 1100°C , a low temperature compared with 1500°C used at industrial scale production. Group II oxides lowered the synthesis temperature of hBN by increasing the boron oxide and ammonia interaction. Also, metal oxides affected the morphology of hBN and G.I.

Acknowledgements: This work was supported by Presidency of Turkey, Presidency of Strategy and Budget, under Contract Number: 2016K12-2834.

REFERENCES

- R. T. Paine, C. K. Narula, "Synthetic routes to boron nitride," *Chem. Rev.*, vol. 90, no. 1, pp. 73 – 91, Jan. 1990.
DOI: 10.1021/cro0099a004
- E. Budak, "Low temperature synthesis of hexagonal boron nitride by solid state reaction in the presence of lithium salts," *Ceram. Int.*, vol. 44, no. 11, pp. 13161 – 13164, Aug. 2018.
DOI: 10.1016/j.ceramint.2018.04.139
- B. Matovic et al., "New synthetic route for nanocrystalline boron nitride powder," *Mater. Lett.*, vol. 65, no. 2, pp. 307 – 309, Jan. 2011.
DOI: 10.1016/j.matlet.2010.10.005
- A. Sumiyoshi, H. Hyodo, K. Kimura, "Li-intercalation into hexagonal boron nitride," *J. Phys. Chem. Solids*, vol. 71, no. 4, pp. 569 – 571, Apr. 2010.
DOI: 10.1016/j.jpcs.2009.12.038
- R. Haubner, M. Wilhelm, R. Weissenbacher, B. Lux, "Boron Nitrides – Properties, Synthesis and Applications" in *High Performance Non-Oxide Ceramics II*, vol. 102, M. Jansen, 1st ed., Berlin/Heidelberg, Germany: Springer, 2002, pp. 1 – 45.
DOI: 10.1007/3-540-45623-6_1
- T. E. O'Connor, "Synthesis of boron nitride," *J. Am. Chem. Soc.*, vol. 84, no. 9, pp. 1753 – 1754, May 1962.
DOI: 10.1021/ja00868a065
- E. Budak, Ç. Bozkurt, "The effect of transition metals on the structure of h-BN intercalation compounds," *J. Solid State Chem.*, vol. 177, no. 4 – 5, pp. 1768 – 1700, Apr. – May 2004.
DOI: 10.1016/j.jssc.2003.12.038
- E. Budak, Ç. Bozkurt, "Synthesis of hexagonal boron nitride with the presence of representative metals," *Phys. B: Condens. Matter*, vol. 405, no. 22, pp. 4702 – 4705, Nov. 2010.
DOI: 10.1016/j.physb.2010.08.067
- M. Hubáček, T. Sato, M. Ueki, "Copper-boron nitride interaction in hot-pressed ceramics," *J. Mater. Res.*, vol. 12, no. 1, pp. 113 – 118, Jan. 1997.
DOI: 10.1557/JMR.1997.0018
- M. Hubáček, T. Sato, T. Ishii, "A coexistence of boron nitride and boric oxide," *J. Solid State Chem.*, vol. 109, no. 2, pp. 384 – 390, Apr. 1994.
DOI: 10.1006/jssc.1994.1117
- L. Chen et al., "Low-temperature synthesis and benzene-thermal growth of nanocrystalline boron nitride," *J. Cryst. Growth*, vol. 273, no. 3 – 4, pp. 646 – 650, Jan. 2005.
DOI: 10.1016/j.jcrysgro.2004.09.062
- J. Thomas, N. E. Weston, T. E. O'Connor, "Turbostratic¹ boron nitride, thermal transformation to ordered-layer-lattice boron nitride," *J. Am. Chem. Soc.*, vol. 84, no. 24, pp. 4619 – 4622, Dec. 1962.
DOI: 10.1021/ja00883a001
- M. G. Balint, M. I. Petrescu, "An attempt to identify the presence of polytype stacking faults in hBN powders by means of X-ray diffraction," *Diam. Relat. Mater.*, vol. 18, no. 9, pp. 1157 – 1162, Sep. 2009.
DOI: 10.1016/j.diamond.2009.02.035
- M. Zheng, Y. Gu, Z. Xu, Y. Liu, "Synthesis and characterization of boron nitride nanoropes," *Mater. Lett.*, vol. 61, no. 8 – 9, pp. 1943 – 1945, Apr. 2007.
DOI: 10.1016/j.matlet.2006.07.108
- G. W. Zhou, Z. Zhang, Z. G. Bai, D. P. Yu, "Catalyst effects on formation of boron nitride nano-tubules synthesized by laser ablation," *Solid State Commun.*, vol. 109, no. 8, pp. 555 – 559, Feb. 1999.
DOI: 10.1016/S0038-1098(98)00541-9
- L. Shi et al., "Formation of nanocrystalline BN with a simple chemical route," *Mater. Lett.*, vol. 58, no. 26, pp. 3301 – 3303, Oct. 2004.
DOI: 10.1016/j.matlet.2004.06.022
- M. Hubáček, T. Sato, "Preparation and properties of a compound in the b-c-n system," *J. Solid State Chem.*, vol. 114, no. 1, pp. 258 – 264, Jan. 1995.
DOI: 10.1006/jssc.1995.1037
- H. E. Çamurlu, A. Gençer, B. Becer, "Comparative catalytic study on the carbothermic formation of hexagonal boron nitride with Li, Na, K and Ca carbonates," *J. Mater. Sci.*, vol. 49, no. 1, pp. 371 – 379, Jan. 2014.
DOI: 10.1007/s10853-013-7714-x